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t ( P ) [13],  
 14, 20 30% DCF t  
 P - t ,  
 [ 15,16] . t  
 ( A P ) DCF -  
 z t [ 19] t 2 [ 17] t [ 18],  
 t 2

$$-\frac{[DCF]}{t} = k [DCF]^m [CO_2]^n \quad (1)$$

$$-\frac{[DCF]}{t} = k_5 [DCF]^m \quad (2)$$

$$k_5 = k [CO_2]^n \quad (3)$$

$$-\frac{[DCF]}{t} = k [DCF]^m [CO_2]^n \quad (4)$$

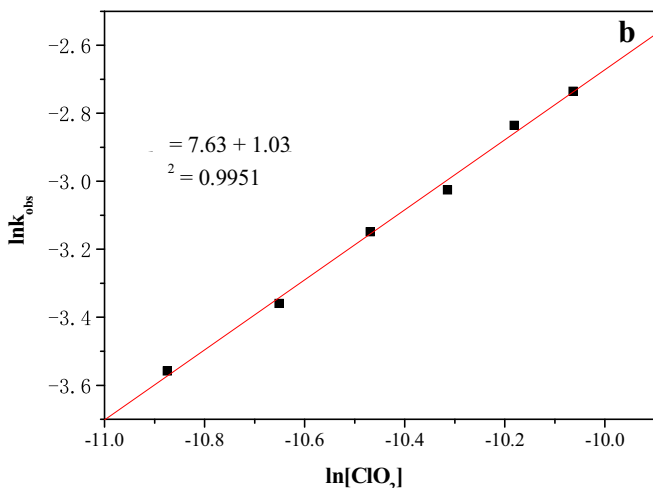
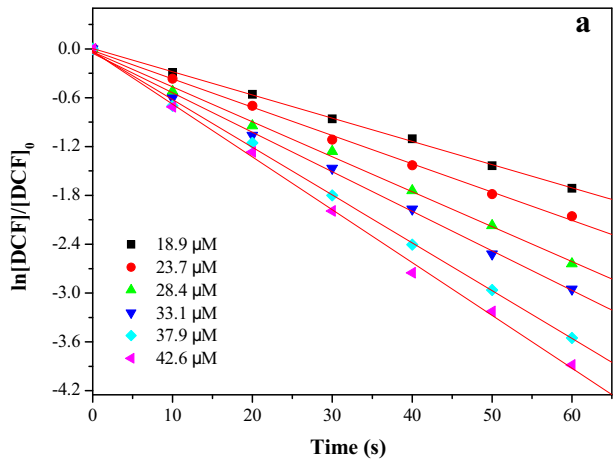


Fig. 1. DCF oxidation by ClO<sub>2</sub> at [DCF]<sub>0</sub> = 3.00 μM, [CO<sub>2</sub>]<sub>0</sub> = 18.9–42.6 μM, T = 7.0 °C.

3.2. Effect of pH on DCF oxidation

$$-\frac{[DCF]}{t} = k_5 [DCF]^m [CO_2]^n \quad (5)$$

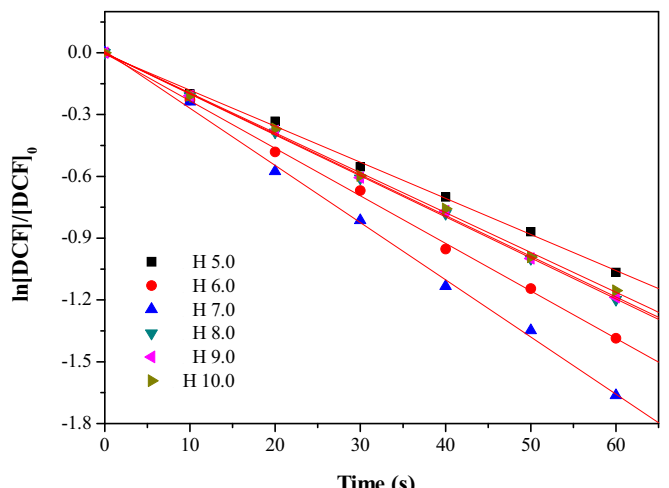


Fig. 2. Effect of pH on DCF oxidation by ClO<sub>2</sub> at [DCF]<sub>0</sub> = 3.00 μM, [CO<sub>2</sub>]<sub>0</sub> = 18.9 μM, T = 7.0 °C.

Table 1

Parameter	Value	Unit
$k_{DCF}$	$1.24(0.022) \times 10^3$	$M^{-1} s^{-1}$
$k_{CO_2}$	$1.72(0.028) \times 10^3$	$M^{-1} s^{-1}$
$A$	278	$s^{-1}$
$A$	308	$s^{-1}$

### 3.3. Effect of temperature on DCF oxidation

The effect of temperature on the rate constant  $k_{DCF}$  was studied. The Arrhenius equation is given by:

$$k = A e^{-\frac{E_a}{RT}} \quad (7)$$

where  $A$  is the pre-exponential factor,  $E_a$  is the activation energy,  $R$  is the gas constant, and  $T$  is the absolute temperature. The values of  $A$  and  $E_a$  were determined from the plot of  $\ln k$  versus  $1/T$ . The values are  $A = 278 s^{-1}$  and  $E_a = 7.87(0.14) \times 10^4 J mol^{-1}$ . The activation energy  $E_a$  is also related to the enthalpy of activation  $\Delta H^\ddagger$  and the entropy of activation  $\Delta S^\ddagger$  by the equation:

$$\ln k = \ln A - \frac{\Delta H^\ddagger}{RT} + \frac{\Delta S^\ddagger}{R} \quad (8)$$

The values of  $\Delta H^\ddagger$  and  $\Delta S^\ddagger$  were determined from the plot of  $\ln k + \frac{\Delta H^\ddagger}{RT}$  versus  $1/T$ . The values are  $\Delta H^\ddagger = 5.38(0.17) \times 10^4 J mol^{-1}$  and  $\Delta S^\ddagger = -165.75(3.83) J mol^{-1} K^{-1}$ .

The equilibrium constant  $K_a$  for the dissociation of DCF is given by:

$$DCF \rightleftharpoons DCF^- + H^+ \quad K_a = 4.2 \quad (5)$$

The equilibrium constant  $K_a$  was determined from the plot of  $\ln K_a$  versus  $1/T$ . The values are  $K_a = 2(0.016) \times 10^3 M^{-1}$  and  $E_a = 2.13(0.032) \times 10^4 J mol^{-1}$ .

The equilibrium constant  $K_a$  for the dissociation of  $CO_2$  is given by:

$$CO_2 + H_2O \rightleftharpoons HCO_3^- + H^+ \quad K_a = 1.96 \quad (6)$$

The equilibrium constant  $K_a$  was determined from the plot of  $\ln K_a$  versus  $1/T$ . The values are  $K_a = 7.5(0.298) \times 10^3 M^{-1}$  and  $E_a = 5.0(7.0) \times 10^4 J mol^{-1}$ .

The values of  $k_B$  and  $E_a$  were determined from the plot of  $\ln k_B$  versus  $1/T$ . The values are  $k_B = 2.63(0.058) \times 10^{-2} s^{-1}$  and  $E_a = 3.51(0.067) \times 10^4 J mol^{-1}$ .

### 3.4. Effect of oxygen concentration on DCF oxidation

The effect of oxygen concentration on the rate constant  $k_{DCF}$  was studied. The rate constant  $k_{DCF}$  was determined from the plot of  $\ln k_{DCF}$  versus  $\ln [O_2]$ . The values are  $k_{DCF} = 2.63(0.058) \times 10^{-2} s^{-1}$  and  $k_{CO_2} = 3.51(0.067) \times 10^{-2} s^{-1}$ .

### 3.5. Kinetics model and mechanism for DCF oxidation

The kinetics model and mechanism for DCF oxidation were studied. The rate constant  $k_{DCF}$  was determined from the plot of  $\ln k_{DCF}$  versus  $1/T$ . The values are  $k_{DCF} = 2.63(0.058) \times 10^{-2} s^{-1}$  and  $k_{CO_2} = 3.51(0.067) \times 10^{-2} s^{-1}$ .

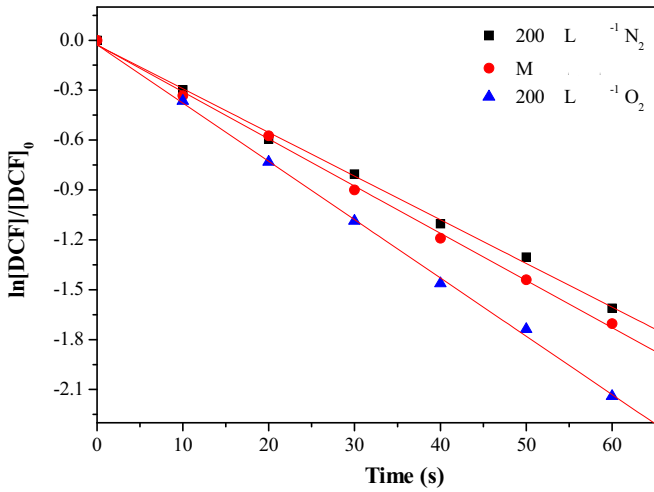


Fig. 4.  $\ln \frac{[DCF]}{[DCF]_0}$  vs Time (s).  $[DCF]_0 = 3.00 \mu M$ ,  $[C \bullet_2]_0 = 7.0 \times 10^{-8} M$ .

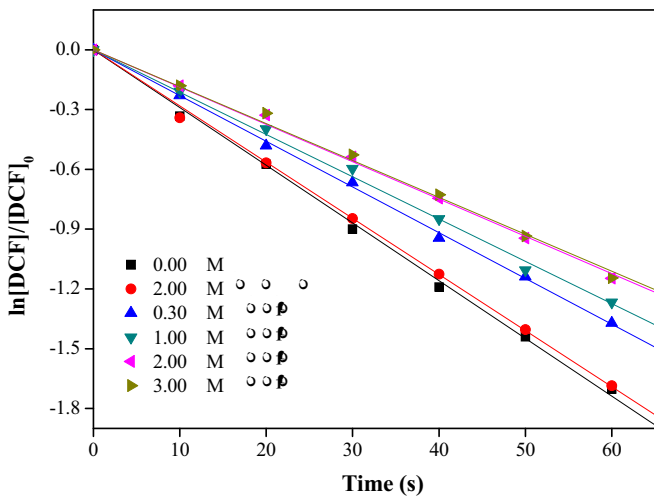
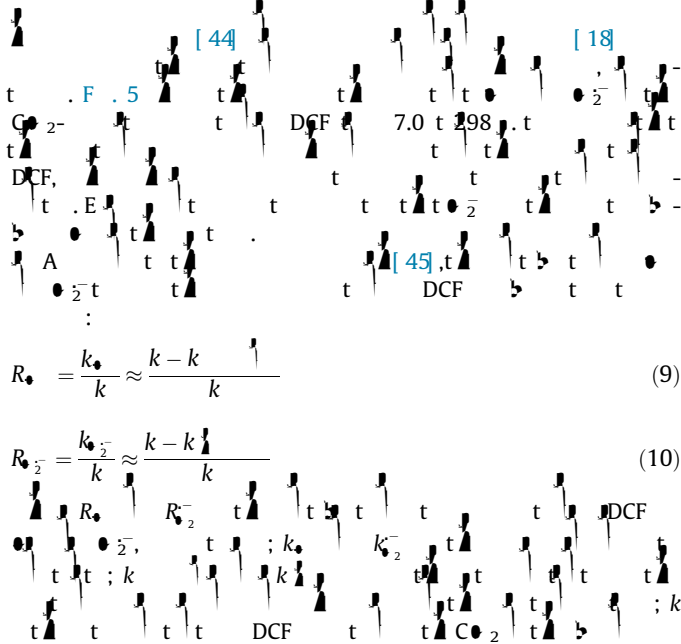
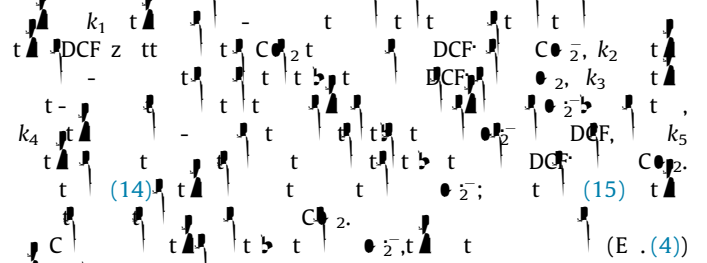
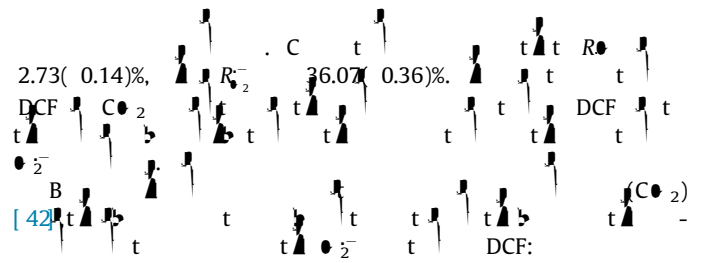


Fig. 5.  $\ln \frac{[DCF]}{[DCF]_0}$  vs Time (s) for different initial concentrations of M.  $[DCF]_0 = 3.00 \mu M$ ,  $[C \bullet_2]_0 = 7.0 \times 10^{-8} M$ .



$$-\frac{[DCF]}{t} = k_1[DCF][C \bullet_2] + k_4[DCF][\bullet_2^-] \quad (16)$$

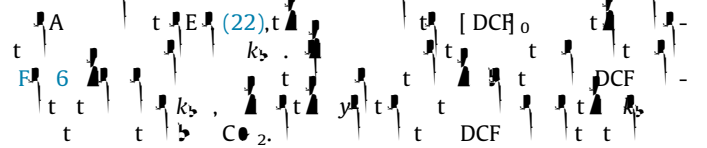
$$\frac{[DCF^+]}{t} = k_1[DCF][C \bullet_2] - k_2[DCF^+][\bullet_2^-] - k_5[DCF^+][C \bullet_2] = 0 \quad (17)$$

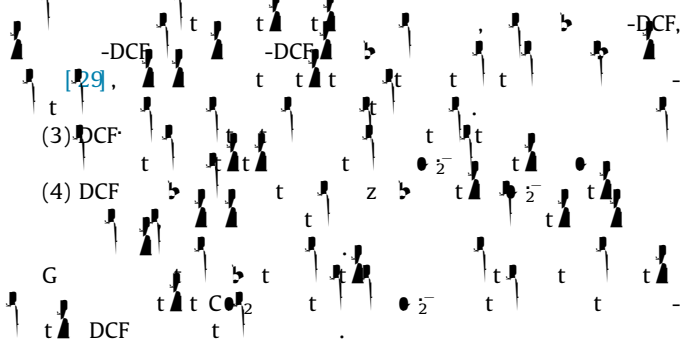
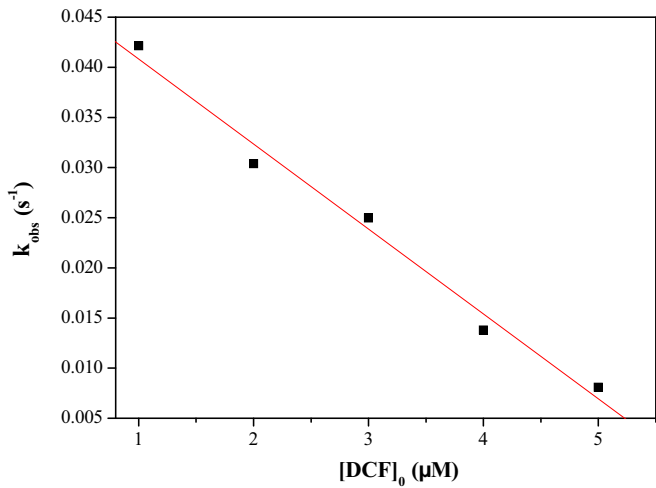
$$-\frac{[\bullet_2^-]}{t} = k_2[DCF^+][\bullet_2^-] - k_3[\bullet_2^-] - k_4[DCF][\bullet_2^-] = 0 \quad (19)$$

$$[\bullet_2^-] = \frac{k_2[DCF^+][\bullet_2^-]}{k_3 + k_4[DCF]} \quad (20)$$

$$-\frac{[DCF]}{t} = \frac{k_1 k_2 k_4 [C \bullet_2] [DCF]^2}{(k_2 [\bullet_2^-] + k_5 [C \bullet_2]) (k_3 + k_4 [DCF])} + k_1 [DCF][C \bullet_2] \quad (21)$$

$$k_5 = -\frac{[DCF]}{t[DCF]} = \frac{k_1 k_2 k_4 [\bullet_2^-] [C \bullet_2]}{(k_2 [\bullet_2^-] + k_5 [C \bullet_2]) (k_3 + k_4 [DCF])} [DCF] + k_1 [C \bullet_2] \quad (22)$$





#### 4. Conclusions

(1) DCF is converted to DCF<sup>+</sup> and DCF-C. The rate constants for these reactions are  $k_1 = 4.21(0.13) \times 10^{-2} s^{-1}$  and  $k_2 = 1.58(0.039) \times 10^{-2} s^{-1}$  respectively.

(2) The rate constant for the reaction of DCF-C with a radical is  $k_3 = 7.87(0.14) \times 10^3 M^{-1} s^{-1}$ .

(3) The overall rate constant for the reaction of DCF with a radical is  $k_4 = 65.75(3.03) \times 10^3 M^{-1} s^{-1}$ .

(4) The overall rate constant for the reaction of DCF with a radical is  $k_5 = 5.38(0.17) \times 10^3 M^{-1} s^{-1}$ .

Fig. 6.  $k_{obs}$  vs  $[DCF]_0$  (18.9  $\mu M$ ). The data points are shown for  $[DCF]_0 = 1.00, 2.00, 3.00, 4.00, 5.00 \mu M$ . The rate constants are  $k_1 = 4.21(0.13) \times 10^{-2} s^{-1}$ ,  $k_2 = 1.58(0.039) \times 10^{-2} s^{-1}$ ,  $k_3 = 7.87(0.14) \times 10^3 M^{-1} s^{-1}$ , and  $k_4 = 65.75(3.03) \times 10^3 M^{-1} s^{-1}$ .

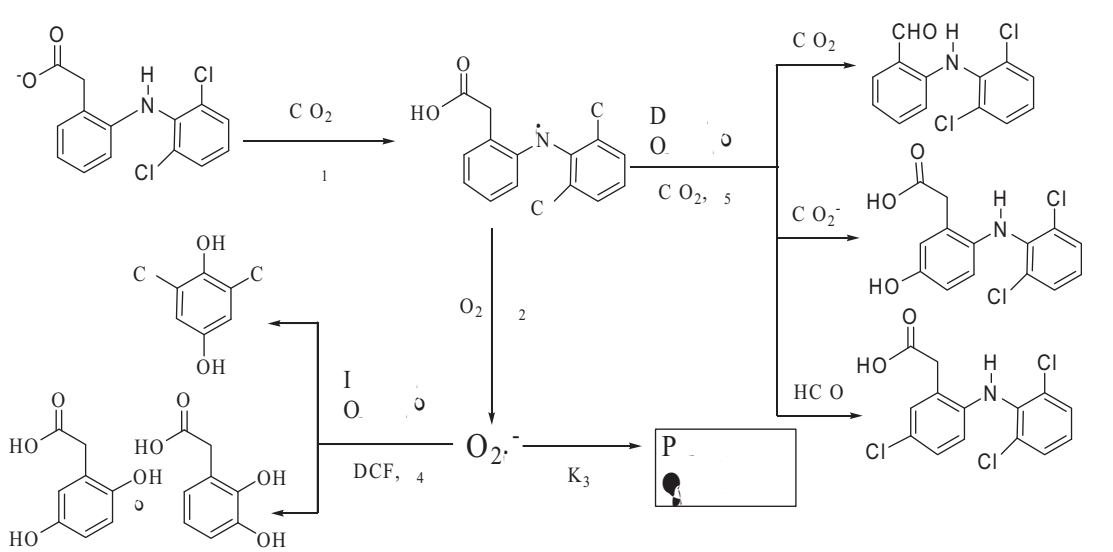
Fig. 7.  $k_{obs}$  vs  $[DCF]_0$  (18.9  $\mu M$ ). The data points are shown for  $[DCF]_0 = 1.00, 2.00, 3.00, 4.00, 5.00 \mu M$ . The rate constants are  $k_1 = 4.96(0.018) \times 10^{-2} s^{-1}$  and  $k_2 = 2.61(0.094) \times 10^{-2} s^{-1}$ .

(1)  $DCF + \bullet C \rightarrow DCF - C + \bullet$

(2)  $DCF + \bullet C \rightarrow DCF + \bullet C$

(3)  $DCF + \bullet C \rightarrow DCF + \bullet C$

(4)  $DCF + \bullet C \rightarrow DCF + \bullet C$



Scheme 1.D

Acknowledgments

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